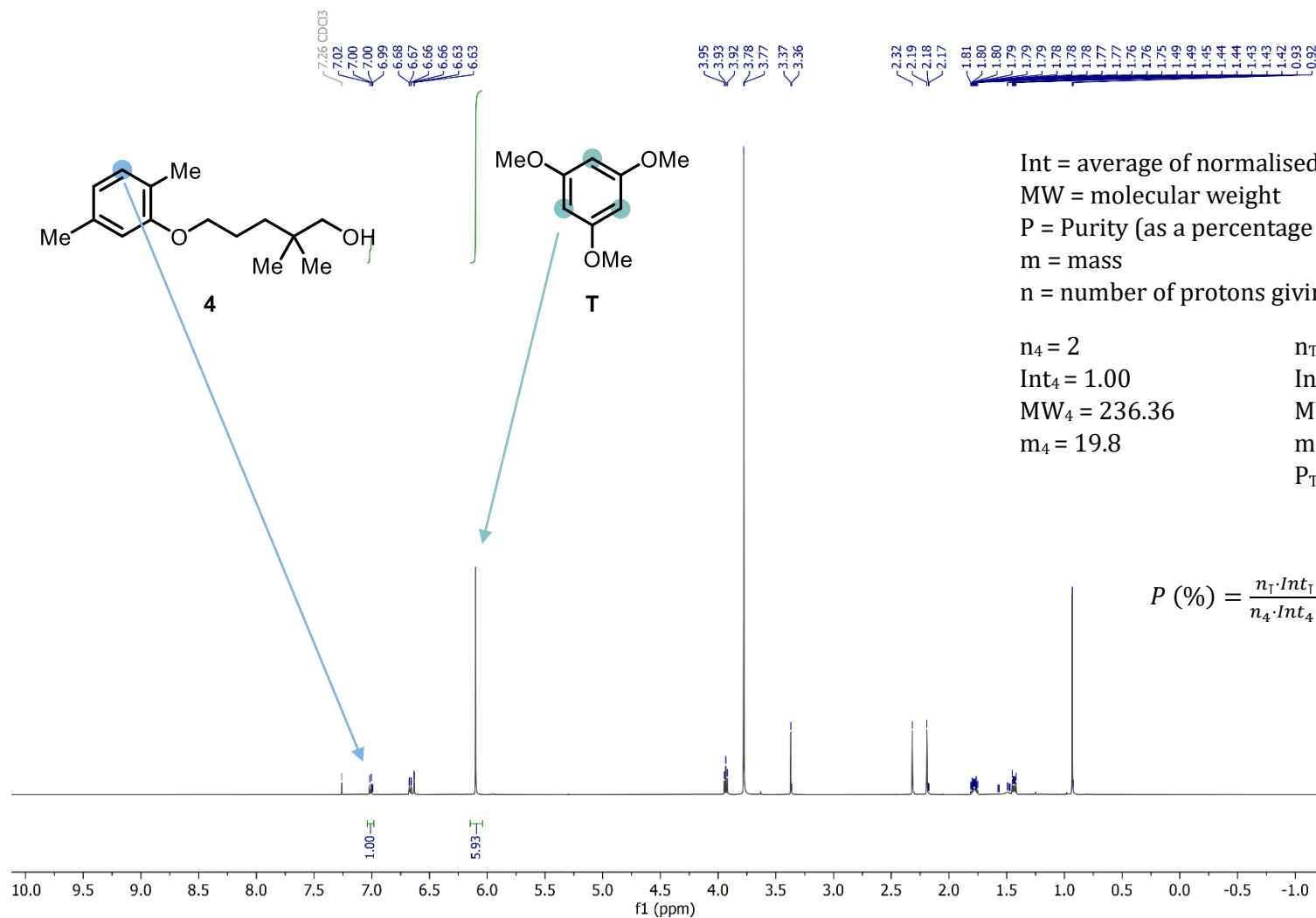


Compound 4 Quantitative ¹H NMR



Int = average of normalised integral values
 MW = molecular weight
 P = Purity (as a percentage value)
 m = mass
 n = number of protons giving rise to a given NMR signal

$n_4 = 2$	$n_T = 3$
$Int_4 = 1.00$	$Int_T = 5.93$
$MW_4 = 236.36$	$MW_T = 168.19$
$m_4 = 19.8$	$m_T = 28.7$
	$P_T = >95\%$

$$P (\%) = \frac{n_T \cdot Int_T \cdot MW_T \cdot m_T}{n_4 \cdot Int_4 \cdot MW_4 \cdot m_4} \cdot P_T = 97.9\%$$